Two types of quasi-liquid layers on ice crystals are formed kinetically

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Surfaces of ice are covered with thin liquid water layers, called quasi-liquid layers (QLLs), even below their melting point (0 °C), which govern a wide variety of phenomena in nature. We recently found that two types of QLL phases appear that exhibit different morphologies (droplets and thin layers) [Sazaki G. et al. (2012) Proc Natl Acad Sci USA 109(4):1052–1055]. However, revealing the thermodynamic stabilities of QLLs remains a longstanding elusive problem. Here we show that both types of QLLs are metastable phases that appear only if the water vapor pressure is higher than a certain critical supersaturation. We directly visualized the QLLs on ice crystal surfaces by advanced optical microscopy, which can detect 0.37-nm-thick elementary steps on ice crystal surfaces. At a certain fixed temperature, as the water vapor pressure decreased, thin-layer QLLs first disappeared, and then droplet QLLs vanished next, although elementary steps of ice crystals were still growing. These results clearly demonstrate that both types of QLLs are kinetically formed, not by the melting of ice surfaces, but by the deposition of supersaturated water vapor on ice surfaces. To our knowledge, this is the first experimental evidence that supersaturation of water vapor plays a crucially important role in the formation of QLLs.

Results and Discussion

We prepared an observation chamber (Fig. S1) in which the temperature of sample ice crystals, $T_{\text{sample}}$, and the partial pressure of water vapor, $P_{\text{H}_2\text{O}}$, can be separately controlled. We kept the total pressure in the chamber at atmospheric pressure using nitrogen gas, and changed the degree of supersaturation of water vapor $\sigma = (P_{\text{H}_2\text{O}} - P_{e})/P_{e}$ by changing $P_{\text{H}_2\text{O}}$ (here $P_{e}$ is the solid–vapor equilibrium pressure). We grew Ih ice single crystals on a cleaved AgI crystal (Fig. S1) at $T_{\text{sample}} = -15.0^\circ\text{C}$ and $P_{\text{H}_2\text{O}} = 585 \text{ Pa (} \sigma = 13\%)$ in a nitrogen environment. After the preparation of ice single crystals, we increased $T_{\text{sample}}$ to a given temperature. All through this process, we kept the sample ice crystals growing, by carefully changing $P_{\text{H}_2\text{O}}$ and confirming the growth by LCM-DIM observations. Then, we observed the behavior of QLLs on basal faces of sample ice crystals under various $P_{\text{H}_2\text{O}}$ ($\sigma$) conditions. LCM-DIM images were processed according to the recipe explained in Fig. S2. Details of the control of $T_{\text{sample}}$ and $P_{\text{H}_2\text{O}}$ were explained in our previous publication (15).

Significance

Thin liquid water layers, so-called “quasi-liquid layers” (QLLs), exist on ice surfaces just below the melting point (0 °C). The formation of QLLs governs various important phenomena on Earth, such as weather- and environment-related issues, winter sports, etc. Hence, QLLs have attracted considerable attention in the fields of ice physics, meteorology, crystal growth, and surface science. Our molecular-level observation of QLLs by advanced optical microscopy reveals that QLLs are formed kinetically as metastable phases only in supersaturated water vapor. This finding opens new horizons of understanding QLLs, which have been so far discussed thermodynamically only at an equilibrium condition.
To reveal the thermodynamic stabilities of the two types of QLL phases, we investigated the water vapor pressure range in which the two types of QLLs could exist on ice basal faces. After we changed $T_{\text{sample}}$ and $P_{\text{H}_2\text{O}}$ from $-2.4 \, ^\circ\text{C}$ and 585 Pa ($\sigma = 13\%$) to $-2.0 \, ^\circ\text{C}$ and 585 Pa ($\sigma = 13\%$), both $\alpha$- and $\beta$-QLLs appeared on a bare ice basal face. Fig. 1 and Movie S1 show the ice basal face after the change. In Fig. 1, white and half-black/white arrowheads represent $\alpha$- and $\beta$-QLL phases, respectively. The black arrowhead (Fig. 1A) depicts an elementary step (0.37 nm in height) growing laterally in the direction of the black arrow (10, 11, 15). In this study, the differential interference contrast was adjusted as if the ice crystal surface was illuminated by a light beam slanted from the upper left to the lower right direction (Fig. S3). Hence, the upper left sides and the lower right sides of convex objects showed brighter and darker contrast, respectively, compared with a flat crystal surface. As marked by white arrowheads in Fig. 1B and C, $\alpha$-QLLs newly appeared on a $\beta$-QLL phase. In addition, adjacent $\beta$-QLLs coalesced with each other (half-black/white arrowheads in Fig. 1B and C), and then $\beta$-QLLs grew in the lateral direction.

Next, we reduced $P_{\text{H}_2\text{O}}$ to 568 Pa (to $\sigma = 10\%$) while keeping $T_{\text{sample}}$ constant at $-2.0 \, ^\circ\text{C}$ and observed how the two types of QLLs behaved by using LCM-DIM (Fig. 2 and Movie S2). The arrowheads and arrows have the same meaning as those in Fig. 1. Elementary steps (black arrowheads) grew laterally in the direction of the black arrow. However, under this condition, the
β-QLL (half-black/white arrowheads) receded gradually. Finally, the β-QLL disappeared completely (Fig. 2 C and D), although α-QLLS (white arrowheads) remained stable.

Then, we further reduced $P_{\text{H}_2\text{O}}$ to 532 Pa (to $\sigma = 3\%$), keeping $T_{\text{sample}}$ constant at $-2.0$ °C, and observed the behavior on the ice basal face. Fig. 3 and Movie S3 show the result. The elementary step marked by the black arrowhead in Fig. 3E grew laterally in the black arrow direction, and then reached the position marked by the black arrowhead in Fig. 3F. Hence, we could confirm that $P_{\text{H}_2\text{O}}$ was definitely in a supersaturated condition ($\sigma > 0$). However, the α-QLL in Fig. 3 (white arrowheads) became gradually smaller as time elapsed. Then, 2 h after Fig. 3A was taken, the α-QLL disappeared completely.

In Figs. 1–3, we examined the critical $P_{\text{H}_2\text{O}}$ above which α- and β-QLLS could appear, keeping $T_{\text{sample}}$ constant at $-2.0$ °C. We performed similar observations at various $T_{\text{sample}}$ to determine the pressure–temperature range in which two types of QLLs could grow. We summarize the results in Fig. 4. Blue filled circles and red open rhombuses show the minimum $P_{\text{H}_2\text{O}}$ at which α- and β-QLLS, respectively, could appear (blue and red solid lines are guides for eyes). In other words, α- and β-QLLS could exist only under $P_{\text{H}_2\text{O}}$ conditions higher than these respective plots. Black open squares present the solid–vapor equilibrium $P_{\text{H}_2\text{O}}$ that we determined experimentally from the LCM-DIM observations of the growth and recession of elementary steps and crystal edges (15). Solid and dashed black curves in Fig. 4 show the solid–vapor equilibrium $P_{\text{H}_2\text{O}}$ and the liquid–vapor equilibrium $P_{\text{H}_2\text{O}}$, respectively, obtained in previous studies (16, 17). The dash-dotted black line also represents the solid–liquid equilibrium $P_{\text{H}_2\text{O}}$ obtained from a calculation using the molar volumes of water and ice, the melting point, and the heat of fusion of ice crystals (18).

Fig. 4 clearly demonstrates that, even just below the melting point, α- and β-QLLS appeared only when $P_{\text{H}_2\text{O}}$ was supersaturated for both ice crystals and liquid water. From this result, we proved, for the first time to our knowledge, that α- and β-QLLS are thermodynamically metastable phases. We also found that β-QLLS could appear only in a higher $P_{\text{H}_2\text{O}}$ range than α-QLLS; β-QLLS are less stable than α-QLLS.

Many researchers have experimentally examined the formation of QLLs under various conditions using different techniques (for details, see Table S1). Some reported that QLLs were formed with $P_{\text{H}_2\text{O}}$ corresponding to an undersaturated condition, and others reported the formation of QLLs under an equilibrium or supersaturated condition. Hence, until we performed the present experiments, there was not a unified view about the role of supersaturation. Moreover, as we reported previously (figure S3 of ref. 15), it is not easy to accurately evaluate the supersaturation of $P_{\text{H}_2\text{O}}$ in the vicinity of a growing or sublimating ice crystal. In our new work, we evaluated the supersaturation of...
$P_{H2O}$ by directly observing the growth and recession of elementary steps. Hence, we believe that we have evaluated the degree of supersaturation more accurately than the other experiments performed so far, and that the results shown in Fig. 4 are substantive.

In the conventional picture of surface melting, it had been thought that only one QLL phase fully covers an ice surface at the solid–vapor equilibrium $P_{H2O}$ (2–4). Lacmann and Stranski first gave a thermodynamic explanation for the wetting of ice crystal surfaces with QLLs (19). Then Kuroda and Lacmann developed this issue into the explanation of morphologies of snowflakes (2). Their theoretical model is based on the competition between the stabilities of a bulk (unstable) and interfaces (stable) of QLLs at the equilibrium $P_{H2O}$ (for details, see Fig. S4) (2). Kuroda and Lacmann evaluated the wettability of a QLL on ice crystal surfaces using values of interfacial free energies determined previously (2, 18, 20), which led them to expect that an ice crystal surface is fully wetted with a QLL (2). Taking into account the contributions of a bulk (unstable) and interfaces (stable) of a QLL, Kuroda and Lacmann assumed that a QLL and an ice can be in equilibrium at the solid–vapor equilibrium $P_{H2O}$, as schematically shown in Fig. 5A. In addition, taking into account the effects of kinetics, they also presumed that a QLL can exist under both undersaturated and supersaturated $P_{H2O}$ conditions (Fig. 5A).

In Fig. 5B, we also summarize the supersaturation range in which the two types of QLLs appeared in our experiments. In contrast to the conventional picture (Fig. 5A), at equilibrium $P_{H2O}$, we found that no QLL phase appeared on ice crystal surfaces. This discrepancy was mainly due to the misestimation of free energy of an ice–vapor interface by Kuroda and Lacmann (2), who estimated the ice–vapor interfacial free energy from the density of dangling bonds on ice crystal surfaces. However, at ultrahigh temperatures (just below the melting point), the effects of entropy significantly decrease the interfacial free energy. Hence, they overestimated the ice–vapor interfacial free energy, leading to the complete wetting of ice crystal surfaces with QLLs (Fig. 5A). Because we can directly observe 2D nucleation on an ice crystal surface by using LCM-DIM (10, 11), we plan to accurately evaluate the ice–vapor interfacial free energy from the measurement of 2D nucleation rates in the future.

In addition to the situation for equilibrium $P_{H2O}$, we found that no QLL phase appeared on ice crystal surfaces under both undersaturated and low supersaturated $P_{H2O}$ conditions (Fig. 5B). QLLs appeared only for $P_{H2O}$ higher than a certain critical supersaturation. This result clearly demonstrates that QLL phases are less stable than ice in any $P_{H2O}$ range. In other words, two types of QLLs are formed kinetically, not by the melting of ice surfaces, but by the deposition of supersaturated water vapor on ice surfaces. To our knowledge, this study shows the first experimental evidence that supersaturation of water vapor plays a crucially important role in the formation of QLLs.

In addition, the two types of QLLs exhibited different stabilities. As demonstrated in this study, β-QLLs appear at a $P_{H2O}$ significantly higher than α-QLLs: the stability of β-QLLs is notably lower than that of α-QLLs. Nevertheless, β-QLLs show a much flatter shape than α-QLLs, implying that β-QLLs have more favorable interaction (i.e., better wettabiliy) with ice surfaces than α-QLLs. To solve this contradiction, we need to obtain more information about the structures of α- and β-QLLs. Hence, we are developing a Raman spectrometer combined with LCM-DIM.

The kinetics of the growth and reduction of α- and β-QLLs is very important. However, to obtain meaningful kinetic data, we need to measure separately the mass transfer processes of water molecules (process 1) between water vapor and QLLs, (process 2) between QLLs and ice surfaces (effects of QLLs on the growth of ice crystals) and (process 3) between α- and β-QLLs. At present, we have no practical idea how to measure the processes 1–3. Hence, the elucidation of these processes is a major challenge for the future.

In this study, we performed in situ LCM-DIM observations on ice basal faces to examine the range of $P_{H2O}$ in which α- and β-QLLs emerged. As a result, we found that α- and β-QLLs are metastable phases that can appear only if $P_{H2O}$ is higher than a critical supersaturation. This result indicates that the two types of QLLs are kinetically formed, not by the melting of ice surfaces, but by the deposition of supersaturated water vapor on ice surfaces. Our findings will provide a key to unlocking the secrets of the various phenomena related with surface melting, from the electrification of thunderclouds to the growth of ice crystals (4).

Materials and Methods

A confocal system (PV500, Olympus Optical Co., Ltd.) was attached to an inverted optical microscope (IX70, Olympus Optical Co., Ltd.), as explained in our recent studies (9, 10). A superluminescent diode (model ASLD68-050-B-FA: 680 nm; Amonics Ltd.) was used as a light source for the LCM-DIM observations. The LCM-DIM system (Fig. S1) used in this study included a large number of the improvements for the molecular-level observation of ice crystal surfaces under atmospheric pressure (10).

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Fig. 5. A comparison between the conventional picture of surface melting (2, 21) (A) and the findings obtained in this study (B). In the conventional picture, one QLL phase appeared irrespective of supersaturation. In contrast, round liquid-like droplets (α-QLLs) and thin liquid-like layers (β-QLLs) appeared only under $P_{H2O}$ conditions higher than the critical saturations.
The observation chamber had upper and lower Cu plates, whose temperatures were separately controlled, using Peltier elements (Fig. S18). At the center of the upper Cu plate, a cleaved AgI crystal (a kind gift from G. Layton of Northern Arizona University), well known as an ice-nucleating agent, was attached using heat grease. On this AgI crystal, the sample ice crystals were grown at ~15°C. To supply water vapor to the sample ice crystals, other ice crystals were grown on the lower Cu plate as a source of water vapor. The temperature of the sample ice crystals, $T_{\text{sample}}$, was determined by controlling the temperature of the upper Cu plate. Because the volume of the source ice crystals on the lower Cu plate was ~10^3 times larger than that of the sample ice crystals, the partial pressure of water vapor $P_{\text{H}_2\text{O}}$ inside the observation chamber was dominated by the equilibrium pressure of the source ice crystals. Hence, $P_{\text{H}_2\text{O}}$ was determined by controlling the temperature of the source ice crystals, $T_{\text{source}}$ (the temperature of the lower Cu plate). Other details, including the calibration of $T_{\text{sample}}$ and $P_{\text{H}_2\text{O}}$, were reported in our recent study (15).

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